

**11<sup>th</sup>**



**International Chemistry Olympiad**

**6 theoretical problems  
2 practical problems**

# THE ELEVENTH INTERNATIONAL CHEMISTRY OLYMPIAD 2-11 JULY 1979, LENINGRAD, SOVIET UNION

## THEORETICAL PROBLEMS

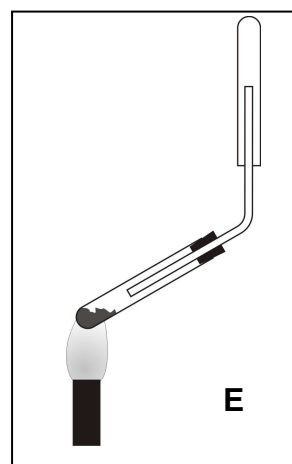
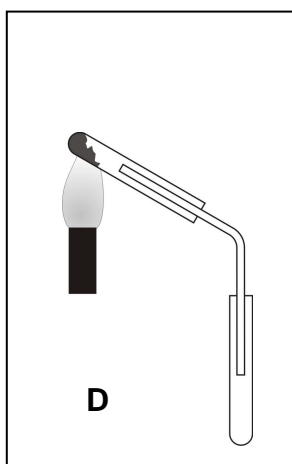
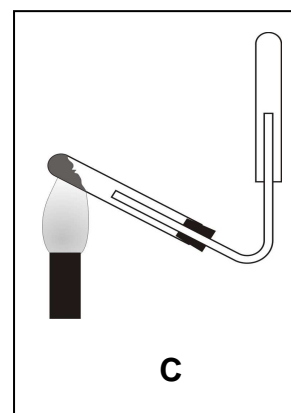
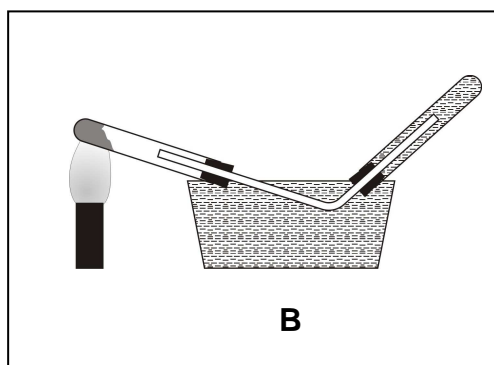
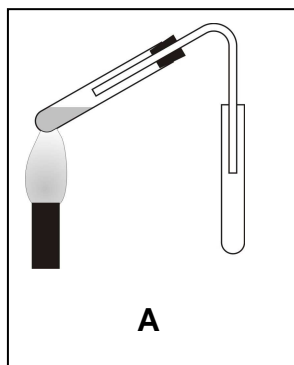
### PROBLEM 1

When carrying out this programmed assignment, encircle those letters which in your opinion correspond to the correct answers to each of the 20 questions.

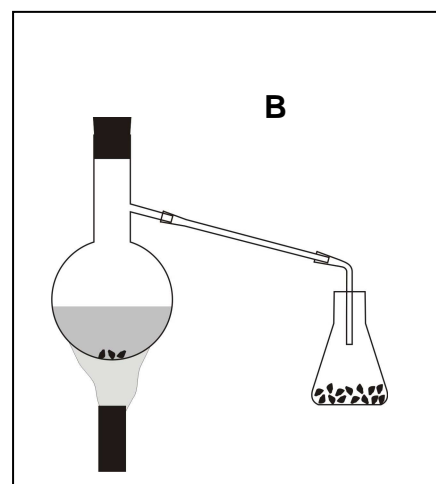
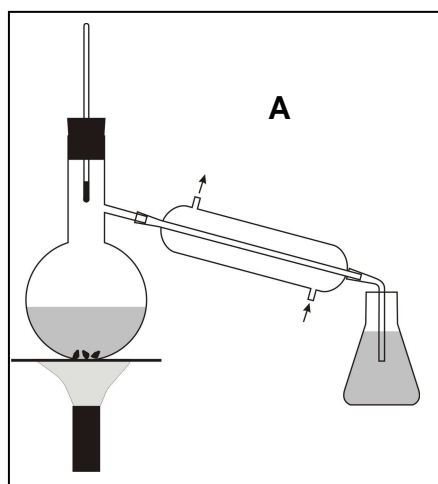
1. Which element is oxidized in the reaction between ethylene and an aqueous solution of potassium permanganate?  
A) carbon, B) hydrogen, C) potassium, D) manganese, E) oxygen.
2. How many litres of CO<sub>2</sub> will approximately be evolved in the reaction of 18 g of potassium hydrogen carbonate with 65 g of 10 % sulphuric acid?  
A) 1, B) 2, C) 3, D) 4, E) 5.
3. Which of the following hydrocarbons gives the maximum heat yield on complete combustion of 1 litre of the gas:  
A) propane, B) methane, C) acetylene, D) ethylene, E) all give the same yield.
4. How many isomers can have a compound if its formula is C<sub>3</sub>H<sub>5</sub>Br?  
A) 1, B) 2, C) 3, D) 4, E) 5.
5. Which of the following hydrocarbons will be the best engine fuel?  
A) cyclooctane, B) 2,2-dimethylhexane, C) normal octane, D) 3-ethylhexane, E) 2,2,4-trimethylpentane.
6. With which of the following compounds will an aqueous solution of a higher oxide of element No 33 react?  
A) CO<sub>2</sub>, B) K<sub>2</sub>SO<sub>4</sub>, C) HCl, D) NaOH, E) magnesium.
7. What must be the minimum concentration (% by mass) of 1 kg of a potassium hydroxide solution for a complete neutralisation of 3.57 moles of nitric acid?  
A) 5 %, B) 10 %, C) 15 %, D) 20 %, E) 25 %.

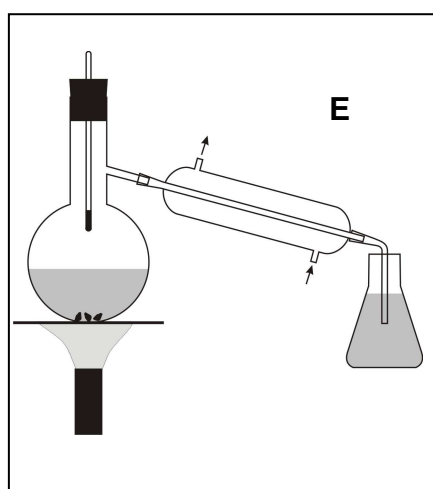
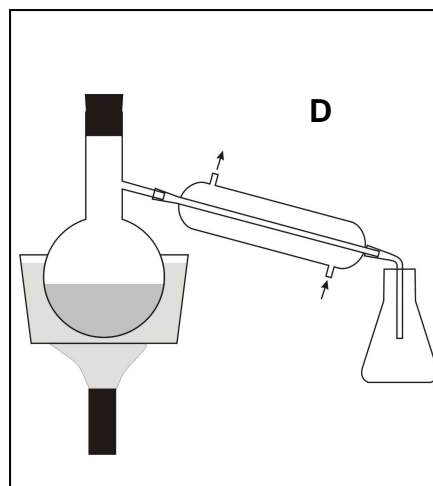
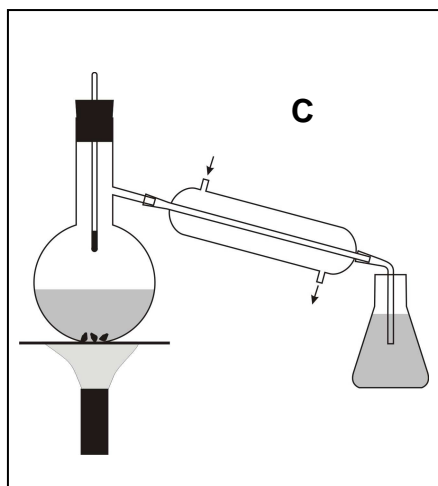
8. How many compounds with the formula  $C_3H_9N$  can exist?  
A) 1, B) 2, C) 3, D) 4, E) 5.
9. In which of the following compounds has the nitrogen content (in mass %) a maximum value?  
A) potassium nitrate, B) barium nitrate, C) aluminium nitrate, D) lithium nitrate, E) sodium nitrate.
10. To which carbon atom (indicate the serial number) will chlorine mainly add in the reaction of HCl with penten-2-oiс acid?  
A) 1, B) 2, C) 3, D) 4, E) 5.
11. How many moles of water are there per mole of calcium nitrate in a crystallohydrate if the water content is 30.5 % by mass?  
A) 1, B) 2, C) 3, D) 4, E) 5.
12. Which of these organic acids is the strongest?  
A) benzoic, B) 2-chlorobenzoic, C) 4-methylbenzoic, D) 2-aminobenzoic, E) 4-bromobenzoic.
13. Which of these acids has the highest degree of dissociation?  
A) HClO, B) HClO<sub>2</sub>, C) HClO<sub>3</sub>, D) HClO<sub>4</sub>, E) all have the same degree.
14. Which of the salts given below do not undergo hydrolysis?  
A) potassium bromide, B) aluminium sulphate, C) sodium carbonate, D) iron(III) nitrate, E) barium sulphate.
15. How many litres of air are approximately required for complete combustion of 1 litre of ammonia?  
A) 1, B) 2, C) 3, D) 4, E) 5.
16. Which element is oxidised in the thermal decomposition of sodium hydrogen carbonate?  
A) sodium, B) hydrogen, C) oxygen, D) carbon, E) none.
17. Which of the following changes have no effect on the chemical equilibrium in the thermal decomposition of CaCO<sub>3</sub>?  
A) temperature elevation, B) pressure decrease, C) addition of catalyst, D) a change in the CO<sub>2</sub> concentration, E) an increase in the amount of the initial substance.
18. Which of the substances given bellow will be formed at the Pt-anode in the electrolysis of an aqueous solution of aluminium chloride?  
A) aluminium, B) oxygen, C) hydrogen, D) aluminium hydroxide, E) chlorine.

19. The apparatus shown in the figures is intended for preparing ammonia under laboratory conditions. The test tube being heated contains a mixture of  $\text{NH}_4\text{Cl}$  and  $\text{Ca}(\text{OH})_2$ . Which of the figures is correct?



20. Which of the apparatuses shown in the figures is the best one for the synthesis of bromethane from potassium bromide, concentrated sulphuric acid and ethanol?





## SOLUTION

- |       |             |              |              |
|-------|-------------|--------------|--------------|
| 1 – A | 6 – D and E | 11 – D       | 16 – E       |
| 2 – C | 7 – D       | 12 – B       | 17 – C and E |
| 3 – A | 8 – D       | 13 – D       | 18 – B and E |
| 4 – E | 9 – D       | 14 – A and E | 19 – C       |
| 5 – E | 10 – C      | 15 – D       | 20 – A       |

**PROBLEM 2**

An alloy comprises the following metals: cadmium, tin, bismuth, and lead. A sample of this alloy weighing 1.2860 g, was treated with a solution of concentrated nitric acid. The individual compound of metal **A** obtained as a precipitate, was separated, thoroughly washed, dried and calcinated. The mass of the precipitate after the calcination to constant mass, was 0.3265 g.

An aqueous ammonia solution was added in excess to the solution obtained after separation of the precipitate. A compound of metal **B** remained in the solution while all the other metals precipitated in the form of sparingly soluble compounds. The solution was first quantitatively separated from the precipitate, and then hydrogen sulphide was passed through the separated solution to saturation. The resulting precipitate containing metal **B** was separated, washed and dried. The mass of the precipitate was 0.6613 g.

The precipitate containing the compounds of metals **C** and **D** was treated with an excess of a NaOH solution. The solution and the precipitate were then quantitatively separated. A solution of HNO<sub>3</sub> was added to the alkaline solution to reach pH 5 – 6, and an excess of K<sub>2</sub>CrO<sub>4</sub> solution was added to the resulting transparent solution. The yellow precipitate was separated, washed and quantitatively transferred to a beaker. Finally a dilute H<sub>2</sub>SO<sub>4</sub> solution and crystalline KI were added. Iodine produced as a result of the reaction was titrated with sodium thiosulphate solution in the presence of starch as an indicator. 18.46 cm<sup>3</sup> of 0.1512 normal Na<sub>2</sub>S<sub>2</sub>O<sub>3</sub> solution were required.

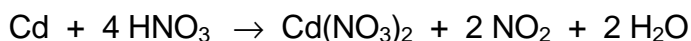
The last metal contained in the precipitate as a sparingly soluble compound was transformed to an even less soluble phosphate and its mass was found to be 0.4675 g.

**2.1** Write all equations of the chemical reactions on which the quantitative analysis of the alloy sample is based. Name metals **A**, **B**, **C**, and **D**. Calculate the mass percentage of the metals in the alloy.

**SOLUTION**

**2.1** The action of nitric acid on the alloy:





Weight form of tin determination:



Calculation of tin content in the alloy:

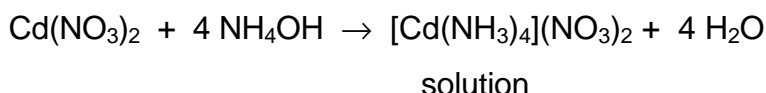
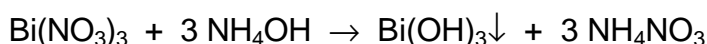
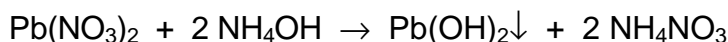
$$M(\text{Sn}) = 118.7 \text{ g mol}^{-1}; \quad M(\text{SnO}_2) = 150.7 \text{ g mol}^{-1}$$

$$\frac{m(\text{Sn})}{m(\text{SnO}_2)} = \frac{M(\text{Sn})}{M(\text{SnO}_2)}; \quad m(\text{Sn}) = \frac{118.7 \text{ g mol}^{-1} \times 0.3265 \text{ g}}{150.7 \text{ g mol}^{-1}} = 0.2571 \text{ g}$$

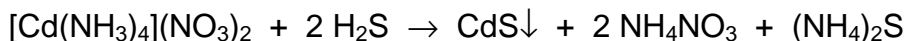
Mass percentage of tin (metal A) in the alloy:

$$w(\text{Sn}) = \frac{0.2571 \text{ g}}{1.2860 \text{ g}} = 0.1999 = 19.99 \%$$

The reactions taking place in the excess of aqueous ammonia solution:



Saturating of the solution with hydrogen sulphide:



Calculation of the cadmium content in the alloy:

$$M(\text{Cd}) = 112.4 \text{ g mol}^{-1}; \quad M(\text{CdS}) = 144.5 \text{ g mol}^{-1}$$

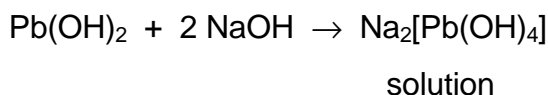
$$m(\text{Cd}) = \frac{112.4 \text{ g mol}^{-1} \times 0.6613 \text{ g}}{144.5 \text{ g mol}^{-1}} = 0.5143 \text{ g}$$

Mass percentage of cadmium (metal B) in the alloy:

$$w(\text{Cd}) = \frac{0.5143 \text{ g}}{1.2860 \text{ g}} = 0.3999 = 39.99 \%$$

The reactions taking place in the excess of sodium hydroxide solution:

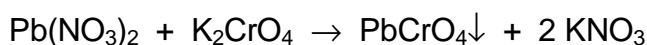
The action of excess sodium hydroxide on lead(II) and bismuth(III) hydroxides:



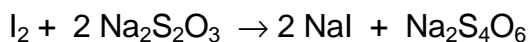
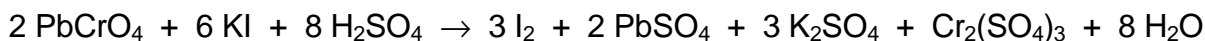
Acidification of the solution with nitric acid (pH = 5 – 6):



The reaction with  $K_2CrO_4$ :



The reactions on which the quantitative determination of lead in  $PbCrO_4$  precipitate is based:



Percentage of lead (metal C) in the alloy:

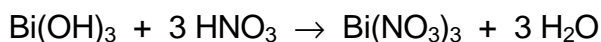
$$w(Pb) = \frac{c(Na_2S_2O_3) \times V(Na_2S_2O_3) \times M(Pb)}{m(\text{alloy}) \times 3}$$

(One  $Pb^{2+}$  ion corresponds to one  $CrO_4^{2-}$  ion which accepts 3 electrons in the redox reaction considered.)

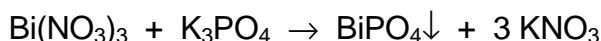
$$w(Pb) = \frac{0.1512 \text{ mol dm}^{-3} \times 0.01846 \text{ dm}^3 \times 207.2 \text{ g mol}^{-1}}{1.286 \text{ g} \times 3} = 0.1499 = 14.99 \%$$

In order to convert bismuth(III) hydroxide to phosphate it is necessary:

a) to dissolve the bismuth(III) hydroxide in an acid:



b) to precipitate  $Bi^{3+}$  ions with phosphate ions:



Calculation of the bismuth content in the alloy:

$$M(Bi) = 209 \text{ g mol}^{-1}; \quad M(BiPO_4) = 304 \text{ g mol}^{-1}$$

$$m(Bi) = \frac{209 \text{ g mol}^{-1} \times 0.4676 \text{ g}}{304 \text{ g mol}^{-1}} = 0.3215 \text{ g}$$

Percentage of bismuth (metal D) in the alloy:

$$w(Bi) = \frac{0.3215 \text{ g}}{1.2860 \text{ g}} = 0.2500 = 25.00 \%$$

Composition of the alloy: % Cd = 40, % Sn = 20, % Pb = 15, % Bi = 25



**PROBLEM 3**

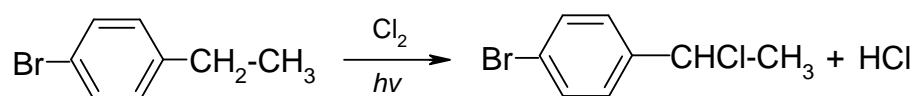
Which chemical processes can take place in the interaction of:

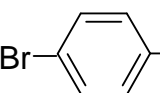
- aluminium ammonium sulphate with baryta water,
- potassium chromate, ferrous chloride and sulphuric acid,
- calcinated soda and sodium hydrogen sulphate,
- 4-bromoethyl benzene and chlorine,
- n-propyl alcohol, phenol and concentrated sulphuric acid?

Write ionic equations for the reactions that proceed in aqueous solutions. For the other chemical reactions write complete equations and indicate the type of the reaction. Indicate the differences in the reaction conditions for those reactions that may lead to the formation of various substances.

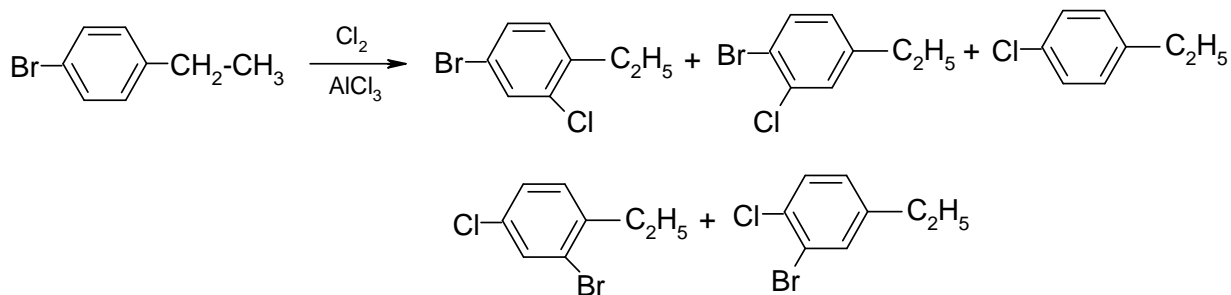
**SOLUTION**

- (a) a-1  $\text{Ba}^{2+} + \text{SO}_4^{2-} \rightarrow \text{BaSO}_4 \downarrow$
- a-2  $\text{NH}_4^+ + \text{OH}^- \rightarrow \text{NH}_3 \cdot \text{H}_2\text{O} \rightarrow \text{NH}_3 \uparrow + \text{H}_2\text{O}$
- a-3  $\text{Al}^{3+} + 3 \text{OH}^- \rightarrow \text{Al}(\text{OH})_3 \downarrow$
- a-4  $\text{Al}(\text{OH})_3 + \text{OH}^- \rightarrow [\text{Al}(\text{OH})_4]^-$
- a-5 possibly:  $\text{Ba}^{2+} + 2 [\text{Al}(\text{OH})_4]^- \rightarrow \text{Ba}[\text{Al}(\text{OH})_4]_2 \downarrow$
- (b) b-1  $2 \text{CrO}_4^{2-} + 2 \text{H}^+ \rightarrow \text{Cr}_2\text{O}_7^{2-} + \text{H}_2\text{O}$
- b-2  $6 \text{Fe}^{2+} + \text{Cr}_2\text{O}_7^{2-} + 14 \text{H}^+ \rightarrow 6 \text{Fe}^{3+} + 2 \text{Cr}^{3+} + 7 \text{H}_2\text{O}$
- b-3 with high concentrations of  $\text{Cl}^-$  and  $\text{H}_2\text{SO}_4$ :  
 $\text{Cr}_2\text{O}_7^{2-} + 4 \text{Cl}^- + 6 \text{H}^+ \rightarrow \text{CrO}_2\text{Cl}_2 + 3 \text{H}_2\text{O}$
- (c) c-1 with excess of  $\text{H}^+$ :  $\text{CO}_3^{2-} + 2 \text{H}^+ \rightarrow \text{H}_2\text{O} \cdot \text{CO}_2 \rightarrow \text{H}_2\text{O} + \text{CO}_2 \uparrow$
- c-2 with excwss of  $\text{CO}_3^{2-}$ :  $\text{CO}_3^{2-} + \text{H}^+ \rightarrow \text{HCO}_3^-$
- (d) d-1 free radical substitution (upon exposure to light or on heating)



small quantity of  and polychlorination

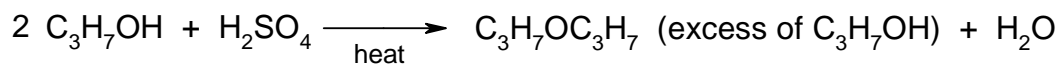
d-2 in the presence of electrophilic substitution catalysts:  
and as side reaction products:



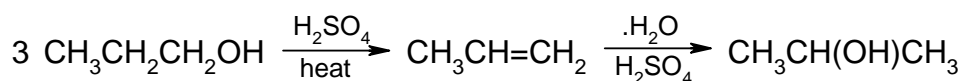
(e) e-1



e-2

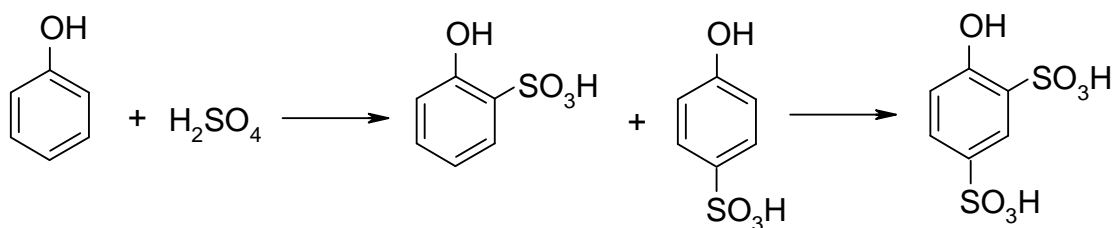


e-3

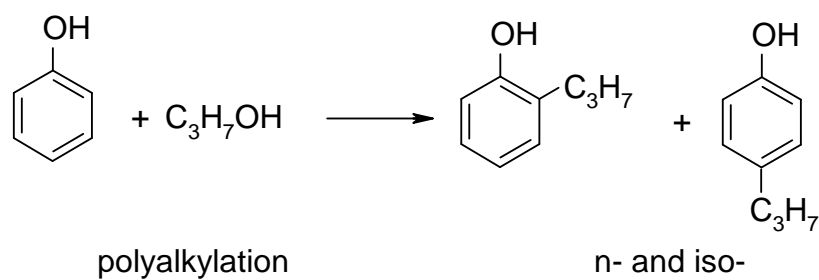


(in e-1 and e-2)

e-4



e-5



e-6 partial oxidation of  $C_3H_7OH$  and  $C_6H_5OH$  with subsequent condensation or esterification

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**PROBLEM 4**

Compound **X** contains nitrogen and hydrogen. Strong heating of 3.2 g of **X** leads to its decomposition without the formation of a solid residue. The resulting mixture of gases is partially absorbed by sulphuric acid (the volume of the gaseous mixture decreased by a factor of 2.8). The non-absorbed gas, that is a mixture of hydrogen and nitrogen, occupies under normal conditions a volume of 1.4 dm<sup>3</sup> and has a density of 0.786 g dm<sup>-3</sup>.

Determine the formula of compound **X**.

**SOLUTION**

If the density of the mixture of N<sub>2</sub> and H<sub>2</sub> is known, its composition can be determined as

$$0.786 \times 22.4 \times (n + 1) = 28n + 2$$

Hence  $n = 1.5$ . The mass of the mixture is  $0.786 \text{ g dm}^{-3} \times 1.4 \approx 1.1 \text{ g}$ . Consequently, the mixture of gases absorbed by sulphuric acid (these gases could be NH<sub>3</sub> and N<sub>2</sub>H<sub>4</sub>) had an average molar mass of

$$\frac{3.2 \text{ g} - 1.1 \text{ g}}{1.4 \text{ dm}^3 \times (2.8 - 1)} \times 22.4 \text{ dm}^3 \text{ mol}^{-1} \approx 18.67 \text{ g mol}^{-1}$$

while NH<sub>3</sub> corresponds to 17 g mol<sup>-1</sup>.

This means that the absorbed gaseous products consist of a mixture of NH<sub>3</sub> and N<sub>2</sub>H<sub>4</sub>.

The composition of the absorbed fraction is

$$\frac{32 + 17n}{n + 1} = 18.67$$

$$n = 8, \text{ i. e. } 8 \text{ NH}_3 + \text{N}_2\text{H}_4.$$

As a result, the overall ratio of the components of the mixture is as follows:

8 NH<sub>3</sub> + N<sub>2</sub>H<sub>4</sub> + 3 N<sub>2</sub> + 2 H<sub>2</sub> which corresponds to a composition of the initial substance **X**: N : H = (2 + 8 + 6) : (4 + 24 + 4) = 1 : 2.

The initial substance **X** is hydrazine N<sub>2</sub>H<sub>4</sub>.

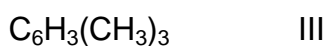
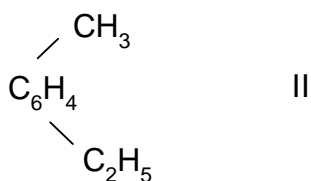
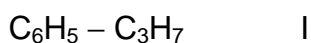
**PROBLEM 5**

Benzene derivative **X** has the empirical formula  $C_9H_{12}$ . Its bromination in the light leads to the formation of two monobromo derivatives in approximately identical yield. Bromination in the dark in the presence of iron also gives two monobromo derivatives. If the reaction is carried out to a higher degree, the formation of four dibromo derivatives may occur.

Suggest the structure for compound **X** and for the bromination products. Write schemes for the reactions.

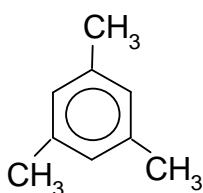
**SOLUTION**

The compound with the empirical formula  $C_9H_{12}$  can be:

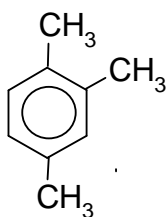


Under the action of bromine in the light without catalysts, bromination of the aliphatic portion will occur, predominantly on the carbon atoms bonded to the aromatic nucleus. When the reaction is conducted in the dark in presence of iron, the latter is converted to  $FeBr_3$  and catalyzes the bromination of the aromatic ring.

Compound **X** cannot be **I** (as then only one monobromo derivative would be formed in the light); it cannot be one of the isomers IIIa, IIIb either.

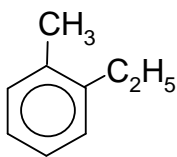


IIIa - Only one monobromo derivative is possible in the bromination of the  $CH_3$  groups.

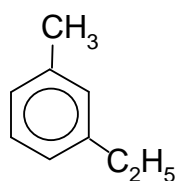


IIIb - Three different monobromo derivatives are possible under the same conditions.

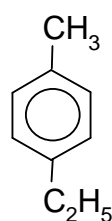
Thus, selection must be made from the following four structures:



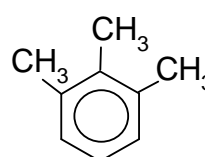
IIa



IIb



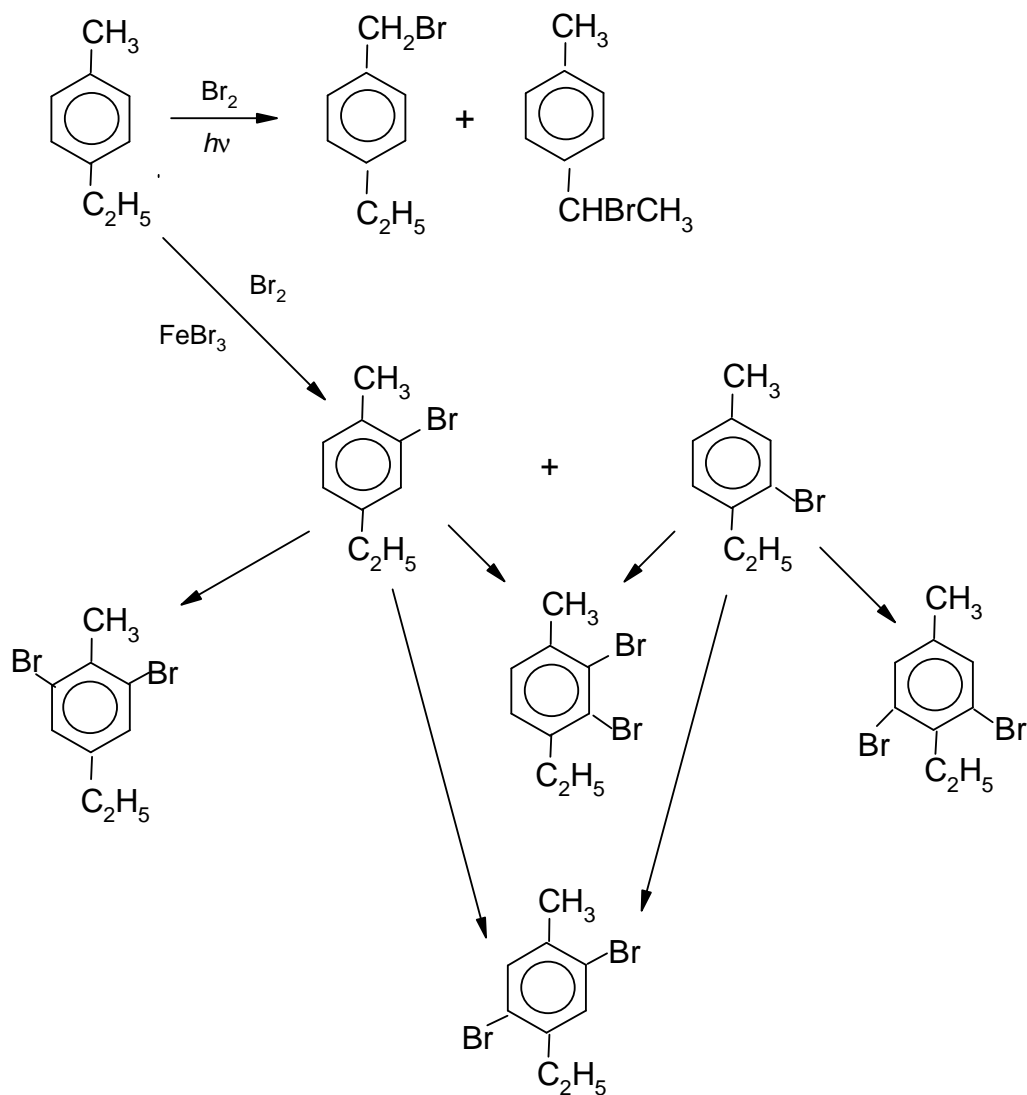
IIc



IIIc

The condition that two monobromo derivatives can be formed in the dark, rules out structures IIa and IIb. The condition of the possibility of four dibromo derivatives rules out structure IIIc. Hence, the only possible structure of compound **X** is IIc.

The scheme of the bromination reaction (next page):



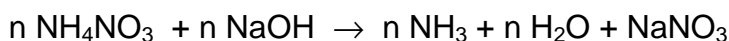
**PROBLEM 6**

130 g of an unknown metal M were treated with excess of a dilute nitric acid. Excess hot alkaline solution was added to the resulting solution and 1.12 dm<sup>3</sup> of a gas evolved (normal conditions).

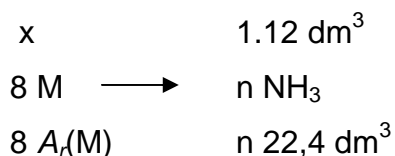
What metal M was dissolved in the nitric solution?

**SOLUTION**

The gas that evolved during the reaction with the alkaline solution was ammonia. Therefore, one of the products resulting from dissolution of the metal M in the acid is ammonium nitrate. Thus, the reaction equations will have the form:



Hence, the scheme:



where  $n$  is the valency of the metal (oxidation number of  $M^{n+}$ ) and  $A_r(M)$  is the relative atomic mass of the metal.

$$8 A_r(M) \Rightarrow 22.4 \times n$$

$$13 \text{ g} \Rightarrow 1.12 \text{ dm}^3$$

$$A_r(M) = \frac{13 \text{ g} \times 22.4 \text{ dm}^3 \times n}{8 \text{ g} \times 1.12 \text{ dm}^3} = 32.5 n$$

If	$n = 1$	then	$A_r(M) = 32.5$	no metal
	$n = 2$		$A_r(M) = 65$	zinc
	$n = 3$		$A_r(M) = 97,5$	none
	$n = 4$		$A_r(M) = 130$	none

Answer: The unknown metal is zinc.



## PRACTICAL PROBLEMS

### PROBLEM 1 (practical)

10 numbered test tubes, 20 cm<sup>3</sup> each, contain 0.1 M solutions of the following substances: barium chloride, sodium sulphate, potassium chloride, magnesium nitrate, sodium orthophosphate, barium hydroxide, lead nitrate, potassium hydroxide, aluminium sulphate, sodium carbonate. Using only these solutions as reagents, determine in which of the numbered test tubes each of the above given substances, is found.

Draw up a plan of the analysis and write equations of the reactions to be carried out. Do not forget to leave at least 2 cm<sup>3</sup> of the solutions in each test tube for checking. If in the course of the analysis an additional quantity of a solution is needed, you may ask the teacher to give it to you but in such case you will lose some points.

### SOLUTION

Table:

	BaCl <sub>2</sub>	Na <sub>2</sub> SO <sub>4</sub>	KCl	Mg(NO <sub>3</sub> ) <sub>2</sub>	Na <sub>3</sub> PO <sub>4</sub>	Ba(OH) <sub>2</sub>	Pb(NO <sub>3</sub> ) <sub>2</sub>	KOH	Al <sub>2</sub> (SO <sub>4</sub> ) <sub>3</sub>	Na <sub>2</sub> CO <sub>3</sub>
BaCl <sub>2</sub>		↓	—	—	↓	—	↓	—	↓	↓
Na <sub>2</sub> SO <sub>4</sub>	↓		—	—	—	↓	↓	—	—	—
KCl	—	—		—	—	—	↓	—	—	—
Mg(NO <sub>3</sub> ) <sub>2</sub>	—	—	—		↓	↓	—	↓	—	↓
Na <sub>3</sub> PO <sub>4</sub>	↓	—	—	↓		↓	↓	—	↓	—
Ba(OH) <sub>2</sub>	—	↓	—	↓	↓		↓	—	↓	↓
Pb(NO <sub>3</sub> ) <sub>2</sub>	↓	↓	↓	—	↓	↓		↓	↓	↓
KOH	—	—	—	↓	—	—	↓		↓	—
Al <sub>2</sub> (SO <sub>4</sub> ) <sub>3</sub>	↓	—	—	—	↓	↓	↓	↓		↓
Na <sub>2</sub> CO <sub>3</sub>	↓	—	—	↓	—	↓	↓	—	↓	

Using the table, the entire problem cannot be solved at once: all the precipitates are white and there are substances that form the same number of precipitates. From the number of precipitates only KCl (1),  $\text{Mg}(\text{NO}_3)_2$  (4), and  $\text{Pb}(\text{NO}_3)_2$  (8) can be determined immediately.

Furthermore,  $\text{Na}_2\text{SO}_4$  and KOH (giving three precipitates each) can be differentiated via the reaction with  $\text{Mg}(\text{NO}_3)_2$  ( $\text{Mg}(\text{OH})_2$ ).

$\text{Ba}(\text{OH})_2$  and  $\text{Al}_2(\text{SO}_4)_3$  (giving 6 precipitates each): through the reaction with KOH ( $\text{Al}(\text{OH})_3$ ).

$\text{BaCl}_2$ ,  $\text{Na}_3\text{PO}_4$  and  $\text{Na}_2\text{CO}_3$  (giving 5 precipitates each): first the reaction with  $\text{Na}_2\text{SO}_4$  indicates  $\text{BaCl}_2$ . Then the reaction with  $\text{BaCl}_2$ :  $\text{Al}_2(\text{SO}_4)_3$  yields  $\text{AlCl}_3$  ( $\text{BaSO}_4$  precipitate is filtered off). Evolution of  $\text{CO}_2$  and formation of  $\text{Al}(\text{OH})_3$  in the reaction with  $\text{AlCl}_3$  solution indicates  $\text{Na}_2\text{CO}_3$ .

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**PROBLEM 2 (practical)**

Determine the mass of potassium permanganate in the solution you are given. You are provided with hydrochloric acid of a given concentration, a potassium hydroxide solution of an unknown concentration, an oxalic acid solution of an unknown concentration, and a sulphuric acid solution (2 N).

Equipment and reagents:

A burette for titration, indicators (methyl orange, litmus, phenolphthalein), pipettes (volumes 10, and 15 or 20 cm<sup>3</sup>), 2 volumetric flasks (250 cm<sup>3</sup>), 2 titration flasks (100 – 150 cm<sup>3</sup>).

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